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1. Write the chemical formulas for the following binary ionic compounds:

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magnesium bromide  $\text{MgBr}_2$  strontium  
iodide  $\text{SrI}_2$  calcium sulfide  $\text{CaS}$   
hydrogen bromide  $\text{HBr}$  lithium oxide  
 $\text{Li}_2\text{O}$  potassium chloride  $\text{KCl}$  ?  
cadmium fluoride  $\text{CdF}_2$  silver sulfide  
 $\text{Ag}_2\text{S}$  potassium phosphide  $\text{K}_3\text{P}$

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A B; lithium fluoride: LiF: lithium chloride: LiCl: lithium bromide: LiBr: lithium iodide: LiI: ...

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Type II Binary Ionic Compounds contain Transition metals (including the Group III, IV, V, VI metals, except

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source:rbpaonline.com In several cases, two elements can combine in a lot of diverse strategies to make different compounds. The very first element receives a prefix whenever there is more than one atom in the mixture.

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Name the following ionic compounds.

Be sure to check for Roman Numerals!

Na<sub>3</sub>P. MgBr<sub>2</sub>. Ag<sub>2</sub>Se. AlCl<sub>3</sub>. Ca<sub>3</sub>P<sub>2</sub>.

Ba<sub>3</sub>N<sub>2</sub>. Li<sub>2</sub>O. ZnF<sub>2</sub>. K<sub>2</sub>S. SrI<sub>2</sub>

(NH<sub>4</sub>)<sub>2</sub>S Ca<sub>3</sub>(PO<sub>3</sub>)<sub>2</sub> Li<sub>2</sub>CO<sub>3</sub>

Al(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>3</sub> KCN AgNO<sub>3</sub> Sb(NO<sub>3</sub>)<sub>5</sub>

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NiS Cu(NO<sub>3</sub>)<sub>2</sub> FeCl<sub>2</sub> Au(CN)<sub>3</sub>  
Pb(SO<sub>4</sub>)<sub>2</sub> Write the correct formulas for each of the ionic compounds.

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Binary ionic compounds always have 2 types of atoms. 3. When naming an ionic compound, you should name the metal first and the nonmetal second

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(metal or nonmetal). 4. Circle one: when naming a binary ionic compound, the ending of the first element does/does not change. 5. When naming a binary ionic compound the nonmetal always has an -ide ending.

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Nomenclature Worksheet 5: Ionic  
Compounds Summary Name the  
following compounds: Give the formula  
for each compound: 23. sodium  
fluoride 24. potassium sulfide ca COQ  
25. calcium carbonate 26. magnesium

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hydroxide (C) 27. zinc nitrate 28. silver acetate 2. 3. 4. 5. 6. 8. 9.  $\text{CaF}_2$   $\text{BaS}$   $\text{CuSO}_4$   $\text{Fe}_2\text{O}_3$   $\text{HgCl}_2$   $\text{AgNO}_3$   $\text{MgCO}_3$   $\text{K}_2\text{HPO}_4$  29. copper (II) oxide

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